

EXPERIMENT 1
ACID-BASE TITRATION

Name : _____ Group : _____

Matric no. : _____ Date of exp. : _____

Lecturer : _____

DATAWeight of Na_2CO_3 = _____ gConcentration of Na_2CO_3 = _____ M**Table 1** : Burette reading in the standardization of HCl solution.

Initial reading (mL)	Final reading (mL)	Volume used (mL)

Average volume of HCl solution = _____ mL

Concentration of HCl solution = _____ M

Table 2 : Burette reading in the standardization of NaOH solution (methyl orange indicator).

Initial reading (mL)	Final reading (mL)	Volume used (mL)

Average volume of NaOH solution = _____ mL

Concentration of NaOH solution = _____ M

Table 3 : Burette reading in the standardization of NaOH solution (phenolphthalein indicator).

Initial reading (mL)	Final reading (mL)	Volume used (mL)

Average volume of NaOH solution = _____ mL

Concentration of NaOH solution = _____ M

Table 4 : Burette reading in the standardization of acid acetic (phenolphthalein indicator).

Initial reading (mL)	Final reading (mL)	Volume used (mL)

Average volume of NaOH solution = _____ mL

Concentration of acid acetic solution = _____ M

Table 5 : Burette reading in the standardization of acetic acid (bromothymol blue indicator).

Initial reading (mL)	Final reading (mL)	Volume used (mL)

Average volume of NaOH solution = _____ mL

Concentration of acetic acid solution = _____ M

Table 6 : Burette reading in the determination of pH curve for strong acid-strong base titration.

pH of solution	Burette reading (mL)	Volume of NaOH solution used (mL)
1.6		
2.2		
3.0		
4.0		
5.0		
6.0		
7.0		
8.0		
9.0		
10.0		
11.0		
11.5		
12.0		

Table 7 : Burette reading in the determination of pH curve for weak acid-strong base titration.

pH of solution	Burette reading (mL)	Volume of NaOH solution used (mL)
2.0		
3.0		
4.0		
5.0		
5.5		
6.0		
7.0		
8.0		
9.0		
10.0		
11.0		
11.5		
12.0		
13.0		